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Part II: Stoichiometry problems 5. If 54.7 grams of propane (C 3 H 8) and 89.6 grams of oxygen (O 2) are available in the balanced combustion reaction to the right: a) Determine which reactant is the limiting reactant. b) Calculate the theoretical yield of CO 2 in grams. 1 mol C 32.00 2 Limiting Reactant: ____ Theoretical Yield: ____

Practice Problems (Chapter 5): Stoichiometry

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Stoichiometry Practice Problems: Practice Problem #1: Oxygen gas can be produced by decomposing potassium chlorate using the reaction below Stoichiometry practice problems #1 answers. If 138. 6 g of KClO 3 is heated and decomposes completely, what mass of oxygen gas is produced? KClO 3 (s) --> KCl (s) + O 2 (g) [unbalanced] Answer to Practice Problem #1 .

Stoichiometry Practice Problems #1 Answers

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